

CHEMISTRY NMDCAT

(UNIT-9)

WWW.SAEEDMDCAT.COM

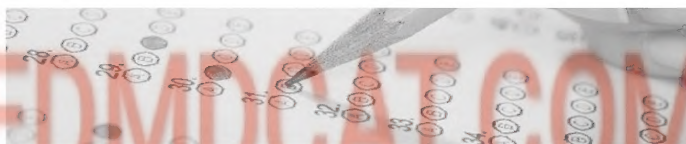
03418729745(WhatsApp Groups)

SAEED MDCAT TEAM

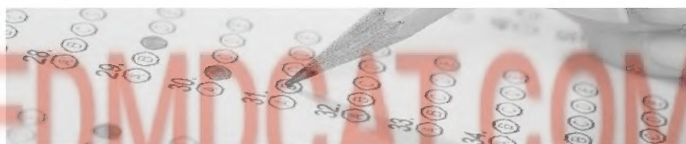
TOPICS:-

✓ s AND p BLOCK ELEMENTS

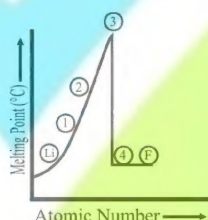
- Q.1** Indicate the parameter, which effect atomic Radii in a period
- Number of shells
 - Nuclear charge
 - Shielding effect
 - No. of orbitals
- Q.2** Size of inert gas is _____ than proceeding halogens
- Larger
 - Smaller
 - Equal
 - Variable
- Q.3** Periodic variation of atomic size in periodic table is similar to that of
- Ionic radii
 - Ionization energy
 - Electron Affinity
 - Electronegativity
- Q.4** Select correct increasing order of atomic radius
- Ne > O > S > Al
 - Ne < O > S > Al
 - Ne < O < S < Al
 - Ne > O < S > Al
- Q.5** Cation is smaller in size than parent atom because of
- Greater effective Nuclear charge and lesser shielding effect
 - Greater effective Nuclear charge and greater shielding effect
 - Lesser effective Nuclear charge and lesser shielding effect
 - Greater effective Nuclear charge and greater shielding effect
- Q.6** Indicate incorrect order of atomic and ionic radii
- Na > Na⁺
 - O⁻² < O⁻¹
 - Cl⁻ > Cl
 - Ar > Cl
- Q.7** When an atom is ionized, the electron is removed from _____.
- Inner most shell
 - Outermost shell
 - Second last shell
 - Any shell
- Q.8** Ionization energy of a neutral atom does not depends upon
- Atomic size
 - Nuclear charge
 - Shielding effect
 - Cationic size
- Q.9** Mark the correct statement, ionization energy of an element in a period increases due to
- Successive addition of e⁻ shell
 - Successive increase of nuclear charge
 - Successive increase of effective nuclear charge
 - Both b and c
- Q.10** With respect to 1st ionization energy, abnormal ionization energy trend are shown by _____ and _____ elements of period 3
- Mg and Al
 - Mg and S
 - Al and S
 - P and S



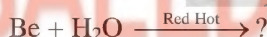
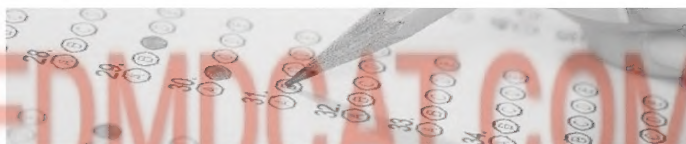
- Q.11** In which of the following configuration removal of e⁻ is easiest
a. $1s^2, 2s^2, 2p^6, 3s^2$
c. $1s^2, 2s^2, 2p^6, 3s^2, 3p^3$
- Q.12** Elements W, X, Y and Z having atomic No 13,14, 15, and 16 respectively which of the following elements will have maximum I.E
a. W
c. Y
- Q.13** In any period of the periodic table, highest ionization energies are shown by
a. Alkali metals
c. Halogens
- Q.14** Which of the following can be electron affinities for the formation of O⁻¹ and O⁻² respectively
a. -141 KJ/mol, +780KJ.mol
c. +141 KJ/mol, -780KJ.mol
- Q.15** Indicate correct set of variation of electron affinities in a period and group respectively
a. Increase and Decrease
c. Decrease and Decrease
- Q.16** Lower value of electron affinity of fluorine is due to
a. Small size
c. Gaseous nature of F
- Q.17** Unit in which electronegativity is measured
a. KJ/mol
c. Joule/sec
- Q.18** Periodic trend of Electronegativity is similar to periodic trend of
a. Ionization energy
c. Atomic size
- Q.19** Which of the following is correct definition of electronegativity
a. Half of the distance b/w nuclei of bounded atoms
b. Amount of energy required to remove electron
c. Capability of an atom to attract shared electrons
d. All of above correctly explains electronegativity
- Q.20** Which of the following generally applies to noble gases
a. High I.E and high reactivity
c. High I.E and low reactivity
- Q.21** Electronegativity of any element can be calculated by comparing with
a. F
c. Br
- Q.22** All the elements which have a tendency to form positive ion by losing electrons are considered as
a. Non-metals
c. Metals
- Q.23** Most metallic element of group IA is
a. Li
c. K
- Q.24** In a period of periodic table change of metallic character is
a. From metals to non-metals
c. From metals to metals
- b. $1s^2, 2s^2, 2p^6, 3s^2, 3p^1$
d. $1s^2, 2s^2, 2p^6, 3s^2, 3p^4$
- b. X
d. Z
- b. Noble gases
d. Alkali earth metals
- b. -141 KJ/mol, -780KJ.mol
d. +141 KJ/mol, +780KJ.mol
- b. Increase and Increase
d. Decrease and Increase
- b. Inter electronic repulsions
d. Both a and b
- b. ev
d. No Unit
- b. Electron affinity
d. Both a and b
- b. Low I.E and low reactivity
d. High E.N and high reactivity
- b. Cl
d. I
- b. Metalloides
d. Gases
- b. Na
d. Rb
- b. From non-metals to metals
d. From non-metals to non-metals



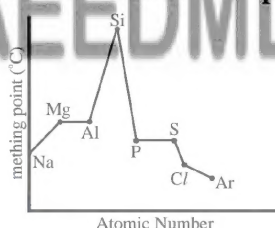
- Q.25 Mark the correct statement**
 a. Non-metallic character decrease as atomic size increases
 b. Non-metals yield acidic oxides
 c. Non-metals are poor conductor of heat and electricity
 d. All of the above
- Q.26 Melting and boiling points of period 2 vary in order**
 a. Increase
 b. Decrease
 c. 1st Increase and then decrease
 d. 1st decrease and then increase
- Q.27 In group VII-A melting points**
 a. Increase down the group
 b. Decrease down the group
 c. 1st Increase and then decrease
 d. 1st decrease and then increase
- Q.28 The diagram below is a plot of melting points of elements of second period against their atomic numbers. Lithium and fluorine are placed at the extreme ends of the plot. On the basis of melting points where would you place carbon among the empty slots on the plot?**

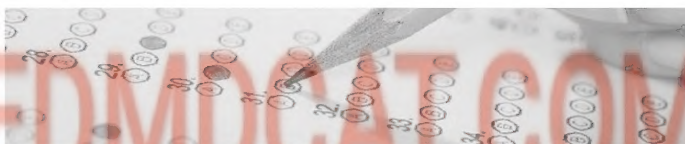


- a. 1
 b. 4
 c. 2
 d. 3
- Q.29 Electrical conductivity is due to free movement of**
 a. Ions
 b. Electrons
 c. Atoms
 d. Nucleus
- Q.30 Electrical conductivity of elements of period 3 increases from Na to _____**
 a. Mg
 b. Al
 c. P
 d. Ar
- Q.31 Lowest melting point is of**
 a. Be
 b. Mg
 c. Ca
 d. Sr
- Q.32 Alkali metals react with H₂O, they produce metal hydroxides and**
 a. CO₂
 b. H₂
 c. O₂
 d. N₂
- Q.33 Li when burn in O₂ mainly gives**
 a. LiO
 b. LiO₂
 c. Li₂O
 d. Li₂O₂
- Q.34 Which of the following alkali metal form peroxide on reacting with O₂**
 a. Na
 b. K
 c. Rb
 d. Cs
- Q.35 $2\text{NaNO}_3 + 10\text{Na} \rightarrow 6\text{Y} + \text{N}_2$, What is "Y" in the given reaction**
 a. Na₂O₂
 b. NaO
 c. Na₂O
 d. NaO₂
- Q.36 Nitrate of which of the following alkali metal decompose to give oxide**
 a. Li
 b. Na
 c. K
 d. Rb
- Q.37 Guess the product formed in given reaction**



- a. $\text{Be}(\text{OH})_2$ b. BeO
c. BeH_2 d. No product is formed
- Q.38 Mark the incorrect statement**
a. Reactivity of IA with H_2O increase down the group
b. Reactivity of IIA with H_2O increase down the group
c. Mg react with H_2O vigorously
d. Calcium reacts with H_2O to form calcium hydroxide
- Q.39 Amphoteric oxide is**
a. Na_2O b. KO_2
c. BeO d. CaO
- Q.40 Metal which gives peroxide on heating with O_2 at high temperature**
a. Be b. Ba
c. Ca d. Mg
- Q.41 Most basic oxide among oxides of group IIA is**
a. BeO b. CaO
c. MgO d. SrO
- Q.42 Cl_2 dissolved in H_2O to form a mixture of HCl and**
a. H_2O b. HOCl
c. HClO_2 d. HClO_3
- Q.43 Aluminum combine with O_2 it does not corrode easily due to formation**
a. $\text{Al}(\text{OH})_3$ b. AlO_2
c. Al_2O_3 d. Al_2O_4
- Q.44 Ionization energy is an index to**
a. Metallic character b. Stability of elements
c. Reactivity of an element d. All of above
- Q.45 Indicate the element which have largest difference between 2nd and 3rd I.E**
a. Na b. Mg
c. Al d. C
- Q.46 In a period moving from Na to carbon increase of melting point is due to**
a. Increase of atomic size b. Decrease of polarizability
c. Increase of binding electrons d. All of above
- Q.47 Maximum electron affinity is possessed by**
a. Cl b. Br
c. I d. F
- Q.48 Only member of group IA that react with N_2 and C is**
a. Na b. K
c. Li d. Cs
- Q.49 The flame colour given by strontium is**
a. Brick red b. Crimson red
c. Violet d. Golden yellow
- Q.50 The trends in melting points of the elements of 3rd period are depicted in figure below**





- a. Decrease in atomic radius from 'Si' to 'P'
- b. Change in bonding and structure of two elements
- c. Different densities of two elements
- d. Increase in electron density from 'Si' to 'P'

**FOR MORE NMDCAT TEST AND
LECTURE**

Visit Our Website: www.saeedmdcat.com

MDCAT WhatsApp Groups: 03418729745

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

Chem T-9

	A	B	C	D		A	B	C	D		A	B	C	D		A	B	C	D
1	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	16	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	31	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	46	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>
2	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	17	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	32	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	47	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
3	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	18	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	33	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	48	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>
4	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	19	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	34	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	49	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>
5	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	20	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	35	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	50	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>
6	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	21	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	36	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	51	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
7	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	22	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	37	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	52	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
8	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	23	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	38	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	53	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
9	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	24	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	39	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	54	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
10	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	25	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	40	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	55	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
11	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	26	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	41	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	56	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
12	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	27	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	42	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	57	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
13	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	28	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	43	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	58	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
14	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	29	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	44	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input checked="" type="radio"/>	59	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>
15	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	30	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	45	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	60	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>

SAEED MDCAT

If you want to get all the test and Lec of all Academy then follow us on different platforms

WEBSITE.www.saeedmdcat.com

FB Group And Page.SAEED MDCAT

INSTAGRAM.SAEED MDCAT

Twitter.Smdcat

For WhatsApp Group.03418729745

Regards.Huzaiifa Saeed,Usama Sohail

Saeed MDCAT

TRUE HERO ALWAYS WINS

National MDCAT 2020 Results

190+ 35 Students

185+ 218 Students

180+ 677 Students

JOIN US FOR FREE
TO SECURE YOUR FUTURE



SAEED MDCAT
SAEED MDCAT TEAM
f SAEEDMDCAT